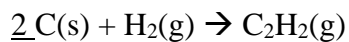
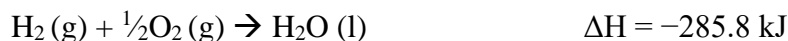
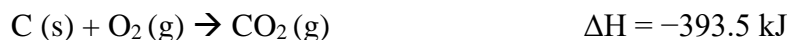
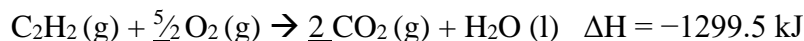


Chem Team Practice - HESS' LAW (#2)

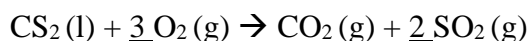
- 1) Calculate the enthalpy (ΔH in kJ) for the following reaction:



given the following thermochemical equations:



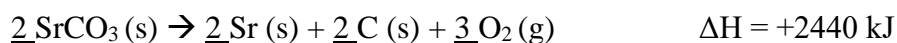
- 2) Calculate the enthalpy of the following chemical reaction:



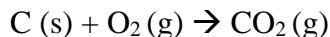
given the following thermochemical equations:



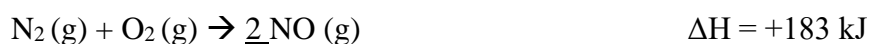
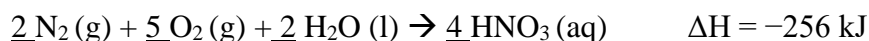
- 3) Given the following data:



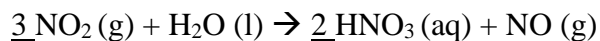
find the ΔH of the following reaction:



- 4) Given the following information:



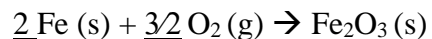
calculate ΔH for the following reaction:



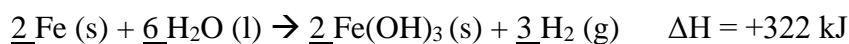
- 5) Calculate ΔH for this reaction: $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + \underline{3}\text{H}_2(\text{g})$
Given the following reactions:



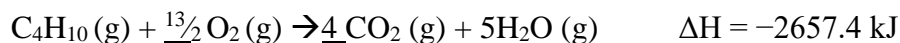
- 6) Determine the heat of reaction for the following reaction:



given the thermochemical equations:



- 7) Determine the standard enthalpy of formation for butane, using the following data:



- 8) Given the equation $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \underline{3}\text{H}_2 + \underline{2}\text{CO}$, calculate delta H.

